**Precipitation/exchange reactions (Level 1) examiners tips: Read these please!**

**•** practise using and be able to interpret the solubility rules chart correctly

 you must link the colour change to the species involved

 learn your colours, both atoms and ions, see help sheet and image below

 *to achieve with Merit you must explain why a precipitate forms...eg "PbCl2 is a precipitate because*

*according to the solubility rules chart, most chlorides are soluble except PbCl2 which is insoluble"*

 *To achieve with Excellence you should refer to the spectator ions in these precipitate reactions*

**Also…”don’t be daft”**

do not write ppt, it is meaningless, write PRECIPITATE, that is the solid substance which forms

precipitates are not soluble, precipitates are **IN**SOLUBLE

You MUST refer to the solubility chart in your Resource book

the symbol of oxygen is a CAPITAL O, this also applies for in nitrates NO3-, carbonates CO32- sulfates SO42-

eh hello....precipitates are NOT insoluble solutions!

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